

Name : \_\_\_\_\_

Team #: \_\_\_\_\_

Name : \_\_\_\_\_

School: \_\_\_\_\_

Event and Construction Requirements

- ☐ 1. Competitors construct only 1 device
- ☐ 2.a. Reference materials secured in 3 ring binder
- ☐ 2.a. Competitors bring up to 2 calculators
- ☐ 2.b. Device, 2 beakers and charts impounded
- ☐ 2.e. Competitors wear eye protection during testing
- ☐ 2.f. Competitors must be able to answer questions about the device per the Building Policy
- ☐ 3.a. No asbestos, mineral wool, fiberglass in device
- ☐ 3.b. Division B: device fits in 20.0 x 20.0 x 20.0cm box
- ☐ 3.c. Division C: device fits in 15.0 x 15.0 x 15.0cm box
- ☐ 3.d. Beaker can be inserted/removed from device
- ☐ 3.e. Device has  $\geq 1.5$ cm diameter hole above beaker
- ☐ 3.e. Top of hole  $< 12$ cm inside beaker bottom
- ☐ 3.e. Hole must remain open during competition
- ☐ 3.f. Device has no energy sources
- ☐ 3.g. Device not significantly different from room temp

**Test Score (back page) (TS)** \_\_\_\_\_

Temperatures (in degrees Celsius)

Start \_\_\_\_\_ Room \_\_\_\_\_

Predicted temperature (PT) \_\_\_\_\_

Final internal beaker (FI) \_\_\_\_\_

Final external beaker (FE) \_\_\_\_\_

Team Verification

4.h. I verify the event supervisor has reviewed with me my team's recorded device testing data

Signature: \_\_\_\_\_

Graphs / Tables (each item can get 0 – 2 points)

Data spans one variable range \_\_\_\_\_

At least 10 data points \_\_\_\_\_

Proper labeling (e.g. title, name, units) \_\_\_\_\_

Number of graphs/tables (x 0.5 points) \_\_\_\_\_

Labeled device picture/diagram \_\_\_\_\_

**Chart Score (CS):** \_\_\_\_\_

Penalties

Competition violations(s)	Y	N
Team didn't bring a device	Y	N
DQed for unsafe operation	Y	N
Construction violations corrected	Y	N
Missed impound	Y	N

Water (ml) \_\_\_\_\_ Ice water (ml) \_\_\_\_\_

**Ice Water Bonus (ml/10) (IWB)** \_\_\_\_\_

Heat Retention Factor (FI/FE) \_\_\_\_\_

Highest Heat Retention Factor (HHRF) \_\_\_\_\_

**Heat Score (HRF/HHRF x 15) (HS)** \_\_\_\_\_

Prediction Est.  $(1 - |(FI - PT)/FI|)$  (PE) \_\_\_\_\_

Highest Prediction Estimate (HPE) \_\_\_\_\_

**Prediction Score (PE/HPE x 25) (PS)** \_\_\_\_\_

**Final Score**  
**(TS+CS+HS+PS+IWB)**

Names : \_\_\_\_\_ School: \_\_\_\_\_ Team #: \_\_\_\_\_

### History of Thermodynamics

Write the corresponding letter next to the date in the column below. (1 pt each, max 10)

- |    |       |      |   |
|----|-------|------|---|
| 1. | _____ | 1650 | A The word “enthapy” is first used to describe energy in a thermodynamic system |
| 2  | _____ | 1697 | B Celsius scale set to 0 (freezing) and 100 (boiling)                           |
| 3  | _____ | 1724 | C William Thomson (Lord Kelvin) defined the need for an absolute zero           |
| 4  | _____ | 1743 | D Otto von Guericke designed and built the world's first vacuum pump            |
| 5  | _____ | 1824 | E Sadi Carnot published “Reflections on the Motive Power of Fire”               |
| 6  | _____ | 1848 | F Thomas Savery built the first engine  |
| 7  | _____ | 1850 | G Rudolf Clausius coined the term “entropy”                                     |
| 8  | _____ | 1876 | H Fahrenheit scale proposed   |
| 9  | _____ | 1909 | I Kelvin scale defined as line from absolute zero and triple point of water     |
| 10 | _____ | 1954 | J Gibbs free energy equation  |

### Temperature Scales and Conversion

11 – 26 In the table below, each row is equal. Fill in the missing numbers. (1 pt each, max 16)

Celsius	Delisle	Fahrenheit	Kelvin	Newton	Rankine	Reaumur	Romer
	559.73		0	-90.14		-218.52	-135.90
-17.78		0	255.37	-5.87	459.67	-14.22	-1.83
-14.29	171.43	6.29	258.864	-4.71		-11.43	
0		32			491.67		7.5
	120	68		6.6	527.67	16	
100	0		373.15		671.67		60

Thermodynamic Laws

Write the corresponding letter next to the date in the column below. (2 pt each, max 16)

Answer	Law	Definition
27 _____	Zeroth law	A the change in the internal energy $\Delta U$ of a closed system is equal to the amount of heat $Q$ supplied to the system, minus the amount of work $W$ done by the system on its surroundings
28 _____	First law	B The state of a simple compressible system is completely specified by two independent, intensive properties
29 _____	Second law	C When the pressure on a sample of a dry gas is held constant, the Kelvin temperature and the volume will be directly related
30 _____	Third law	D the total entropy can never decrease over time for an isolated system
31 _____	Boyle's law	E <i>If A is in thermal equilibrium with B and if B is in thermal equilibrium with C, then A is in thermal equilibrium with C</i>
32 _____	Charles' law	F The entropy of a perfect crystal at absolute zero is exactly equal to zero
33 _____	Dalton's law	G a mixture of non-reacting gases, the total pressure exerted is equal to the sum of the partial pressures of the individual gases
34 _____	State Postulate	H For a fixed amount of an ideal gas kept at a fixed temperature, pressure and volume are inversely proportional

Thermodynamics Concepts

Circle the most appropriate answer for each question (2 pts each, 12 max)

35. Which of the following is NOT a mode for heat transfer:

- |                     |                      |
|---------------------|----------------------|
| a. Conduction       | b. Radiation         |
| c. Retention        | d. Convection        |
| e. All of the above | f. None of the above |

36. Entropy can be best be defined as::

- |   |                                      |
|---|--------------------------------------|
| a. Measure of heat capacity of a system   | b. Degree of disorder in a system    |
| c. Amount of heat transferred by a system | d. Extent a system is open or closed |
| e. All of the above                       | f. None of the above                 |

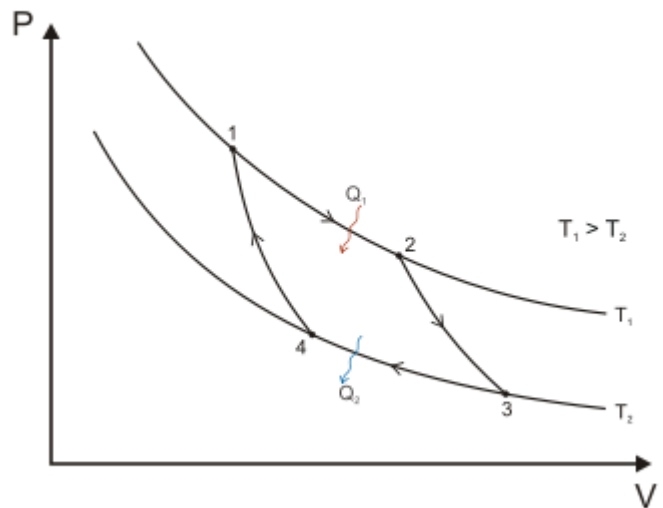
37. The ideal gas law implies:

- a.  $PV/T = \text{constant}$
- b. Initial  $(PV/T) = \text{Final } (PV/T)$
- c.  $PV = nRT$
- d.  $R = PV/T$  divided by number of moles
- e. All of the above
- f. None of the above

38. Heat capacity divided by unit mass is equal to:

- a. Latent heat
- b. Enthalpy
- c. Energy
- d. Conductivity
- e. All of the above
- f. None of the above

Questions 39 & 40 pertain to the graph at the right.



39. This graph represents:

- a. Brayton cycle
- b. Carnot cycle
- c. Ericsson cycle
- d. Rankine cycle
- e. All of the above
- f. None of the above

40. The line segment 2-3 is best described as:

- a. Isothermal Expansion
- b. Adiabatic Expansion
- c. Isothermal Compression
- d. Adiabatic Compression
- e. All of the above
- f. None of the above

### Thermodynamic Problems

Answer the following problems in the space provided. Show your work for potential partial credit. (46 pts max)

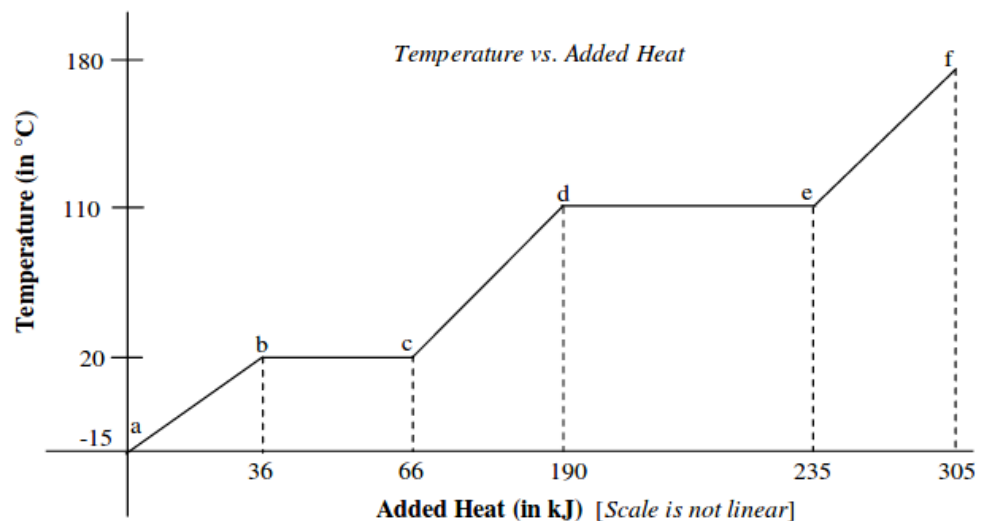
41. A heat engine rejects waste heat to a nearby lake at a rate of 20kW. The lake's temperature is 12C, and the outside temperature is 20C. The engine is operating on a Carnot cycle with a thermal efficiency of 80%. Determine the power output of the engine and the temperature of the source, in C. (10 pts)

42. A mass of 15 kg gains 98.0 kJ of heat when its temperature increases 32F to 68F. Determine the specific heat of the mass. (10 pts)

43. A piston-cylinder device has a volume of 2 liters at 15C and 1 atmosphere. The pressure inside increases by 10 kPa and the temperature by 5C. Determine the new volume inside the piston-cylinder. (10 pts)

44. Insulation with a conductivity of  $0.5 \text{ W/m}\cdot\text{K}$  has a thickness of  $25\text{mm}$ . The insulation is replaced with  $30\text{mm}$  thick insulation that has a conductivity of  $0.45 \text{ W/m}\cdot\text{K}$ . Will the new insulation be more effective (save energy) or be less effective (lose energy)? Explain answer and determine by how much (% better or worse)? (6 pts)

For questions 45 & 46, consider the following. A  $10\text{-g}$  material is studied to understand its phase changes. The graph below shows the transition from gas to solid.



45. At what temperature does the material melt? (4 pts)

46. What is the latent heat of vaporization? (6 pts)

47. How much heat is required to raise the substance from 20C to 110C. Explain your answer. (6 pts)

48. A 5 kg rock is heated to 200C and then placed into a 20C bucket of water. If the bucket contains 50 kg of water, what temperature will the water be once it reaches equilibrium? Assume the rock has a specific heat of 0.8 kJ / (kg K). (6 pts)

Score (S):

Highest Score (HS):

Test Score (S/HS x 45):